

Perhalogenated pyrimidine scaffolds. Reactions of 5-chloro-2,4,6-trifluoropyrimidine with nitrogen centred nucleophiles

Emma L. Parks¹, Graham Sandford^{*1}, John A. Christopher²
and David D. Miller²

Full Research Paper

Open Access

Address:

¹Department of Chemistry, University of Durham, South Road, Durham, DH1 3LE, U.K. and ²GlaxoSmithKline R&D, Medicines Research Centre, Gunnels Wood Road, Stevenage, Hertfordshire SG1 2NY, U.K.

Email:

Emma L. Parks - e.l.parks@durham.ac.uk; Graham Sandford* - graham.sandford@durham.ac.uk; John A. Christopher - john.a.christopher@gsk.com; David D. Miller - david.d.miller@gsk.com

* Corresponding author

Keywords:

pyrimidine; rapid analogue synthesis; perfluoroheteroaromatic; nucleophilic aromatic substitution

Beilstein Journal of Organic Chemistry **2008**, 4, No. 22.

doi:10.3762/bjoc.4.22

Received: 19 March 2008

Accepted: 20 June 2008

Published: 01 July 2008

© 2008 Parks et al; licensee Beilstein-Institut.

License and terms: see end of document.

Abstract

Background

Highly functionalised pyrimidine derivatives are of great importance to the life-science industries and there exists a need for efficient synthetic methodology that allows the synthesis of polysubstituted pyrimidine derivatives that are regioselective in all stages to meet the demands of RAS techniques for applications in parallel synthesis. 5-Chloro-2,4,6-trifluoropyrimidine may be used as a scaffold for the synthesis of polyfunctional pyrimidine systems if sequential nucleophilic aromatic substitution processes are regioselective.

Results

Use of 5-chloro-2,4,6-trifluoropyrimidine as a core scaffold for the synthesis of functionalised pyrimidine systems is assessed in reactions with a small range of nitrogen centred nucleophiles. Mixtures of products arising from nucleophilic aromatic substitution processes are formed, reflecting the activating effect of ring nitrogen and the steric influences of the chlorine atom.

Conclusions

5-Chloro-2,4,6-trifluoropyrimidine is not an ideal scaffold for analogue synthesis or for multiple substitution processes because purification must be performed to remove the 2-substituted regioisomer from the mixture before further reactions can be carried out. However, 4-amino derivatives can be isolated in acceptable yields using this methodology.

Introduction

Highly functionalised pyrimidine derivatives are of great importance to the life-science industries and, indeed, many pyrimidine derivatives have been used for various medicinal applications (Figure 1) [1-3].

Synthesis of pyrimidine rings most commonly involves cyclocondensation reactions of amidine, guanidine or thiourea derivatives with either 1,3-diketone or 1,3-diester systems [4,5]. However, many of these reactions are not regioselective and, furthermore, there is an added difficulty of synthesizing a range of structurally related pyrimidine analogues by parallel synthesis or rapid analogue synthesis (RAS) techniques [6,7] due to the limited range of non-cyclic polyfunctional precursors available. These limitations have, in part, provided added impetus for drug discovery programmes to develop effective synthetic methodology towards multiply substituted systems from simple readily accessible pyrimidine scaffolds [8]. Consequently, pyrimidine core scaffolds that bear multiple functionality, which may be transformed into a widely diverse range of functionalised derivatives by a sequence of efficient and regioselective reactions, are becoming increasingly important [6,7]. In particular, the attachment of amino groups to the pyrimidine nucleus by formation of carbon-nitrogen bonds is a highly desirable process but one of the most difficult to achieve in practice.

Pyrimidines are electron-deficient aromatic systems and, when halogenated, become very useful substrates for a variety of nucleophilic aromatic substitution (S_NAr) processes [9] and, since numerous chloropyrimidines are commercially available, there have been many reports of synthetic strategies concerned with creating pyrimidine-based libraries from halogenated core scaffolds. For example, recently, synthesis of an inhibitor of the

cyclin-dependent kinase was developed [10] (Scheme 1) starting from 2,4,6-trichloropyrimidine as the core scaffold. However, as regioisomeric products are formed in both nucleophilic aromatic substitution stages, separation of the isomers is required after each step, making adoption of this scaffold for analogue synthesis less likely.

There remains, therefore, a requirement for efficient synthetic methodology that allows the synthesis of polysubstituted pyrimidine derivatives that are regioselective in all sequential nucleophilic aromatic substitution stages to meet the demands of RAS techniques for applications in parallel synthesis.

We are exploring the use of polyhalogenated heteroaromatic systems [11-13] as potential hetaryl core scaffolds for analogue synthesis of polyfunctional heterocyclic systems [14-16]. Polyhaloaromatic systems act as useful scaffolds because, in principle, several or all halogen atoms can be displaced by nucleophiles, giving rise to a wide range of heteroaromatic systems and, in this context, we have used a range of perfluorinated heteroaromatic molecules as synthetically versatile building blocks for the creation of new molecular scaffolds for drug discovery [14-17]. In this paper, we describe the reactivity of 5-chloro-2,4,6-trifluoropyrimidine (**1**) with a range of representative nitrogen centred nucleophiles, with the aim of exploring the regioselectivity of these nucleophilic aromatic substitution processes in order to assess the utility of the system as a scaffold for pyrimidine analogue synthesis. Whilst **1** is commercially available and has been known for some time, only a limited number of reactions have been reported (e.g. with ammonia to give the 4-amino derivative [18]) despite the fact that **1** is used widely in the fibre reactive dye industry [19]. However, a systematic study of the reactivity of this potentially

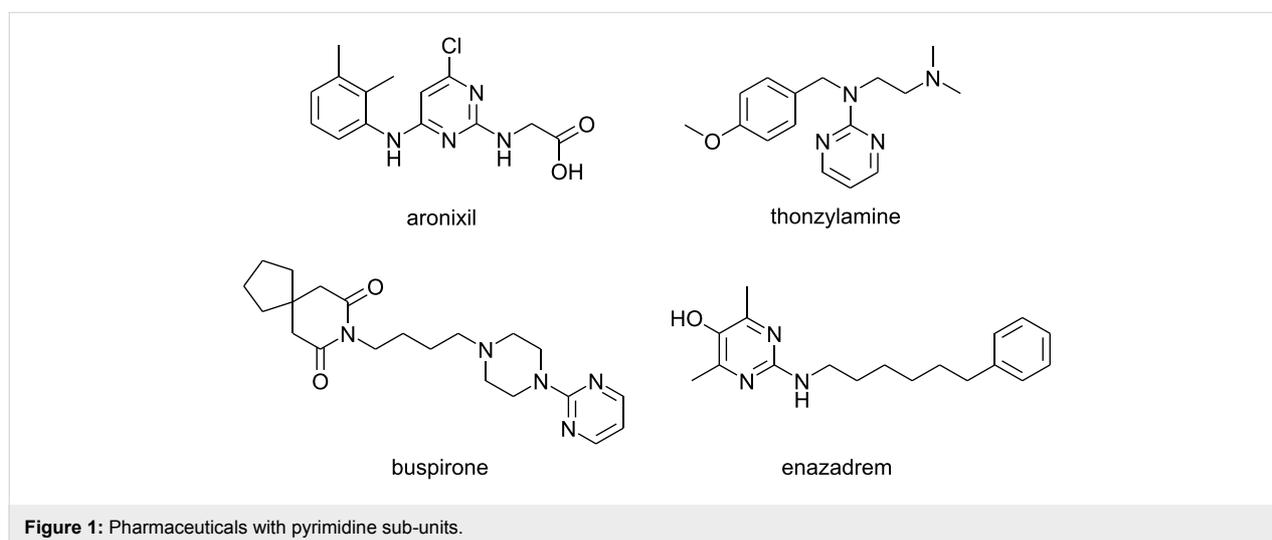


Figure 1: Pharmaceuticals with pyrimidine sub-units.

valuable scaffold with other nitrogen nucleophiles has not been reported.

Results and Discussion

A series of reactions between 5-chloro-2,4,6-trifluoropyrimidine (**1**) and a range of primary and secondary amines were carried out in acetonitrile at 0 °C in the presence of DIPEA as a hydrogen fluoride scavenger and these results are collated in Table 1. All of the reactions were monitored via ¹⁹F NMR and the isomer ratios measured by ¹⁹F NMR integration from samples taken from the reaction mixture.

Reaction of **1** with ammonia results in two isomeric products, as observed by ¹⁹F NMR analysis of the reaction mixture which displayed two distinctive peaks (−48.18 and −69.47 ppm) for the 4-substituted isomer and one peak (−65.44 ppm) for the 2-isomer in a 9 : 1 ratio, the chemical shifts being consistent with previous studies [18]. Similarly, reaction of **1** with ethylamine gives two isomers in an 8 : 1 ratio by ¹⁹F NMR as shown by the appearance of two fluorine signals (−47.48 and −70.83 ppm) and one signal (−63.59 ppm) corresponding to the 4- and 2-amino isomers respectively. Distillation afforded the 4-isomer in good yield. Other reactions gave a mixture of products which were identified by ¹⁹F NMR as described above and, in all cases, the major product could be isolated by either recrystallisation, or column chromatography. All products were fully characterised and ¹⁹F NMR analysis of the crude reaction mixtures gave the ratio of products observed.

Furthermore, when **1** was reacted with the difunctional nucleophile benzamidine, nucleophilic substitution of the fluorine at the 4- and the 2-position in a 40 : 1 ratio occurred (Scheme 2). The main product **3g** was isolated by recrystallisation from acetonitrile and characterised by X-ray crystallography (Figure 2).

In all cases, therefore, the major product obtained arises from substitution of the fluorine atom at the 4-position which is the most activated site *para* to ring nitrogen and further activated by the adjacent chlorine atom, consistent with previous observations for reactions involving perfluorinated heterocycles [13]. However, as the steric requirement of the nucleophile increases, the amount of product arising from substitution at the less activated 2-position is increased, reflecting the steric hindrance of substitution at the more activated 4-position by the larger chlorine atom located at the adjacent 5-position.

Reaction of the related 2,4,6-trifluoropyrimidine with ammonia is reported [20] to give two products in a 4 : 1 ratio and a primary amine, ethanolamine, gave a 2 : 1 ratio of products. Therefore, reactions of **1** with nitrogen centred nucleophiles are more selective than 2,4,6-trifluoropyrimidine despite the increased steric influence of the chlorine atom to nucleophilic attack. This can be rationalised by the fact that the electronegative chlorine atom activates the site *ortho* to itself towards nucleophilic attack and this partly compensates for steric factors in these reactions.

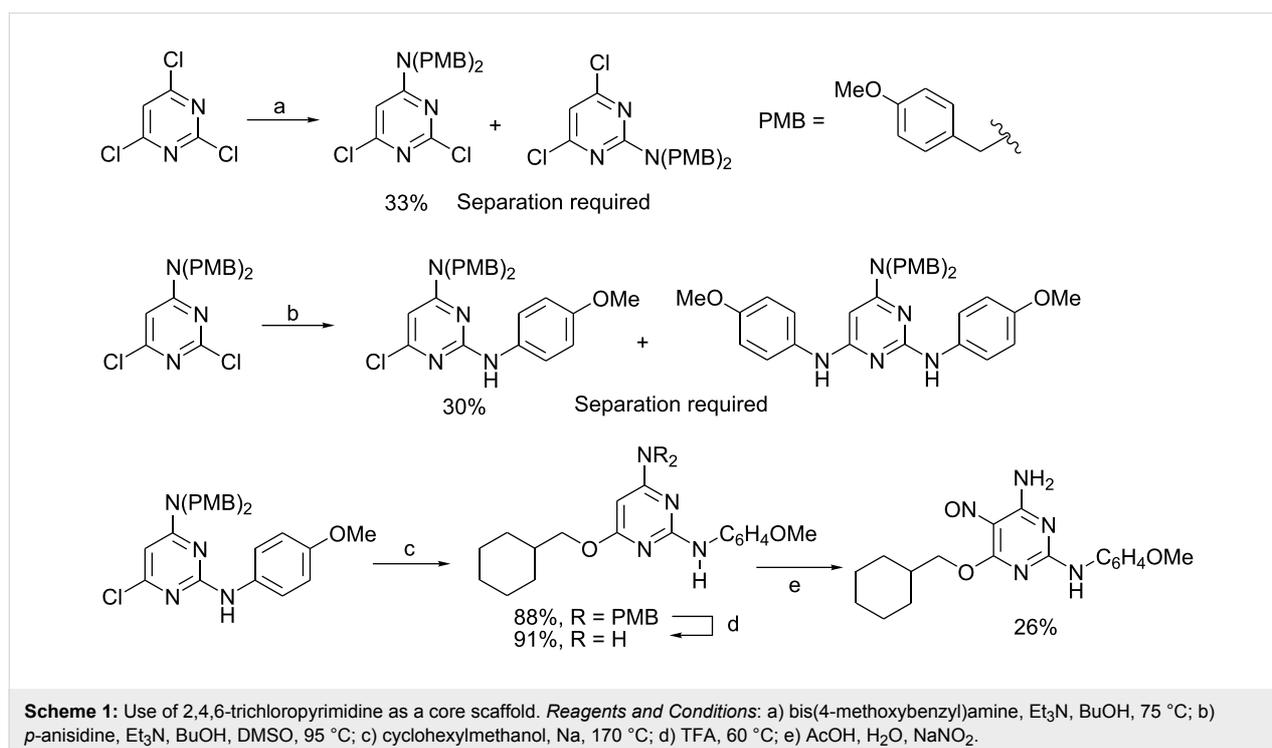
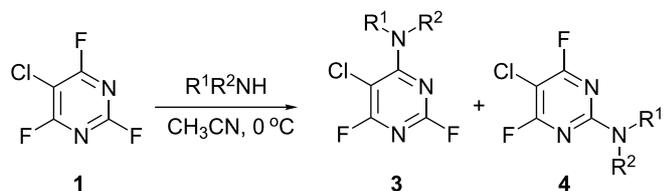
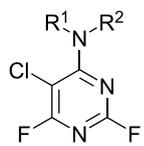
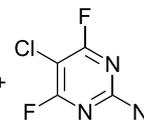
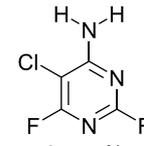
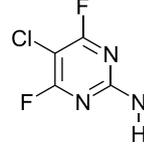
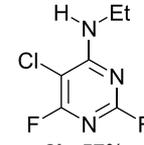
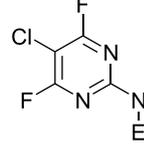
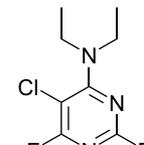
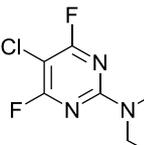
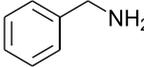
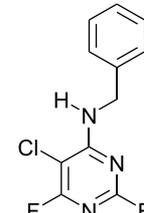
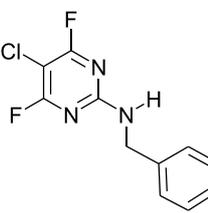
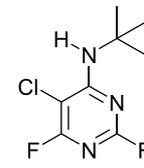
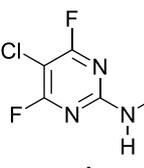
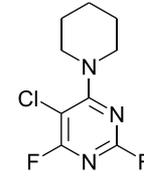
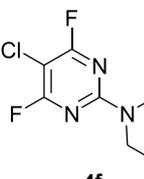
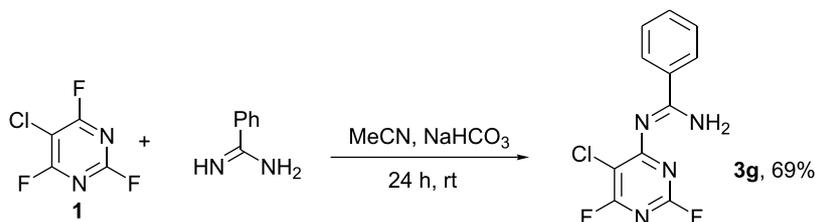


Table 1: Reactions of amine nucleophiles with 5-chloro-2,4,6-trifluoropyrimidine (**1**).

	R^1R^2NH	Products (isolated yield) ^a	Ratio 3 : 4 ^b	
				
				
		 3a , 57%	 4a	9 : 1
		 3b , 57%	 4b	8 : 1
		 3c , 47%	 4c	5 : 1
		 3d , 41 %	 4d	5 : 1
		 3e , 49%	 4e	3 : 1
		 3f , 49%	 4f	3 : 1

^a Isolated yield of major products **3**. Minor products **4** were not isolated.^b Ratio of **3** : **4** in crude product mixture by ¹⁹F NMR analysis.



Scheme 2: Reaction of 1 with benzamidine.

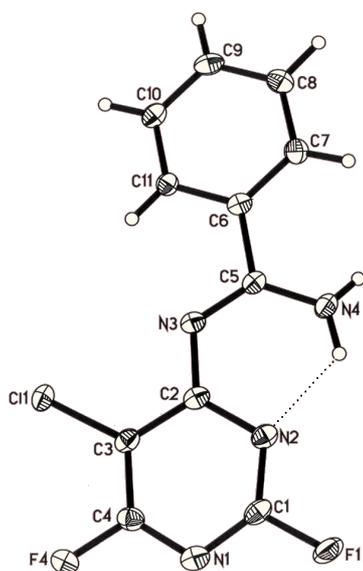


Figure 2: Molecular structure of 3g.

Consequently, it becomes clear that 5-chloro-2,4,6-trifluoropyrimidine (**1**) is not an ideal scaffold for analogue synthesis or for multiple substitution processes because purification must be performed to remove the 2-substituted regioisomer from the mixture before further reactions can be carried out. However, 4-amino derivatives can be isolated in acceptable yields using this methodology and, indeed, these systems could be used as scaffolds for further analogue synthesis.

Experimental

Typical Procedure: Synthesis of *N*-Benzyl-5-chloro-2,6-difluoropyrimidin-4-amine (**3d**)

A solution of 5-chloro-2,4,6-trifluoropyrimidine (0.5 g, 3 mmol), benzylamine (0.32 g, 3 mmol) and DIPEA (0.39 g, 3 mmol) in acetonitrile (50 cm³) was stirred at 0 °C for 2 h after which time ¹⁹F NMR indicated 100% conversion with the formation of *N*-benzyl-5-chloro-2,6-difluoropyrimidin-4-amine (**3d**) (−45.80 and −67.84 ppm) and *N*-benzyl-5-chloro-4,6-

difluoropyrimidin-2-amine (**4d**) (−48.09 ppm) in a 5 : 1 ratio. The reaction solvent was evaporated and the crude product partitioned between DCM (3 × 40 cm³) and water (40 cm³). The organic layer was separated, dried (MgSO₄) and evaporated to dryness to give a crude product containing **3d** and **4d** as a yellow solid (0.54 g). Recrystallisation from *n*-hexane yielded *N*-benzyl-5-chloro-2,6-difluoropyrimidin-4-amine (**3d**) (0.31 g, 41%) as a white solid; mp 57–59 °C; IR (neat, ν, cm^{−1}): 3408, 3281, 2364, 2169, 1739, 1612, 1528, 1447, 1349, 1129, 695; (Found: C, 51.7; H, 3.1; N, 16.6; C₁₁H₈ClF₂N₃ requires: C, 51.7; H, 3.15; N, 16.4%); δ_H (CDCl₃) 4.74 (2H, d, ²J_{HH} 5.8, CH₂), 7.39 (5H, m, Ar-H); δ_C (CDCl₃) 46.2 (s, CH₂), 93.1 (dd, ²J_{CF} 21.4, ⁴J_{CF} 8.0, C-5), 128.1 (s, Ar-CH), 128.4 (s, Ar-CH), 129.2 (s, Ar-CH), 136.9 (s, Ar-CH), 159.3 (dd, ¹J_{CF} 222, ³J_{CF} 22.1, C-2), 162.6 (dd, ³J_{CF} 13, ³J_{CF} 5.4, C-4), 164.5 (dd, ¹J_{CF} 236.2, ³J_{CF} 18.7, C-6); δ_F (CDCl₃) −45.8 (1F, s, C-6), −67.9 (1F, s, C-2); *m/z* (EI⁺) 255 ([M]⁺, 40%), 218 (10), 178 (12).

All other experimental procedures and data are presented in Supporting Information File 1 which accompanies this paper.

Supporting Information

Supporting Information File 1

Experimental procedures and data.

[<http://www.beilstein-journals.org/bjoc/content/supplementary/1860-5397-4-22-S1.doc>]

Acknowledgments

We thank EPSRC for funding (DTA studentship to ELP), GSK Pharmaceuticals for their support of this work and Prof. Judith Howard and Dr Dmitrii Yufit for X-ray crystallography.

References

- Katritzky, A. R.; Rees, C. W., Eds. *Comprehensive Heterocyclic Chemistry*; Pergamon Press: Oxford, 1984; Vol. 3.
- Luo, G.; Chen, L.; Poindexter, G. S. *Tetrahedron Lett.* **2002**, *43*, 5739–5742. doi:10.1016/S0040-4039(02)01190-5
- Pozharskii, A. F.; Soldatenkov, A. T.; Katritzky, A. R. *Heterocycles in Life and Society*; John Wiley and Sons: New York, 1997.

4. Joule, J. A.; Mills, K. *Heterocyclic Chemistry*, 4th ed.; Blackwell: Oxford, 2000.
5. Eicher, T.; Hauptmann, S. *The Chemistry of Heterocycles*; Thieme: Stuttgart, 1995.
6. Collins, I. J. *Chem. Soc., Perkin Trans. 1* **2000**, 2845–2861. doi:10.1039/a904715h
7. Collins, I. J. *Chem. Soc., Perkin Trans. 1* **2002**, 1921–1940. doi:10.1039/b108580h
8. Gordon, E. M.; Kerwin, J. F., Eds. *Combinatorial Chemistry and Molecular Diversity in Drug Discovery*; John Wiley and Sons: New York, 1998.
9. Iddon, B.; Suschitzky, H. Polychloroheteroaromatic compounds. In *Polychloroaromatic Compounds*; Suschitzky, H., Ed.; Plenum Press: London, 1974; pp 197 ff.
10. Marchetti, F.; Sayle, K. L.; Bentley, J.; Clegg, W.; Curtin, N. J.; Endicott, J. A.; Golding, B. T.; Griffin, R. J.; Haggerty, K.; Harrington, R. W.; Mesguiche, V.; Newell, D. R.; Noble, M. E. M.; Parsons, R. J.; Pratt, D. J.; Wang, L. Z.; Hardcastle, I. R. *Org. Biomol. Chem.* **2007**, *5*, 1577–1585. doi:10.1039/b703241b
11. Chambers, R. D.; Sargent, C. R. *Adv. Heterocycl. Chem.* **1981**, *28*, 1–71.
12. Brooke, G. M. *J. Fluorine Chem.* **1997**, *86*, 1–76. doi:10.1016/S0022-1139(97)00006-7
13. Chambers, R. D. *Fluorine in Organic Chemistry*, 2nd ed.; Blackwell: Oxford, 2004.
14. Sandford, G.; Slater, R.; Yufit, D. S.; Howard, J. A. K.; Vong, A. *J. Org. Chem.* **2005**, *70*, 7208–7216. doi:10.1021/jo0508696
15. Baron, A.; Sandford, G.; Slater, R.; Yufit, D. S.; Howard, J. A. K.; Vong, A. *J. Org. Chem.* **2005**, *70*, 9377–9381. doi:10.1021/jo051453v
16. Hargreaves, C. A.; Sandford, G.; Slater, R.; Yufit, D. S.; Howard, J. A. K.; Vong, A. *Tetrahedron* **2007**, *63*, 5204–5211. doi:10.1016/j.tet.2007.03.164
17. Cartwright, M. W.; Sandford, G.; Bousbaa, J.; Yufit, D. S.; Howard, J. A. K.; Christopher, J. A.; Miller, D. D. *Tetrahedron* **2007**, *63*, 7027–7035. doi:10.1016/j.tet.2007.05.016
18. Banks, R. E.; Field, D. S.; Haszeldine, R. N. *J. Chem. Soc. C* **1967**, 1822–1826. doi:10.1039/J39670001822
19. Banks, R. E.; Smart, B. E.; Tatlow, J. C., Eds. *Organofluorine Chemistry. Principles and Commercial Applications*; Plenum: New York, 1994.
20. Delia, T. J.; Anderson, D. P.; Schomaker, J. M. *J. Heterocycl. Chem.* **2004**, *41*, 991–993.

License and Terms

This is an Open Access article under the terms of the Creative Commons Attribution License (<http://creativecommons.org/licenses/by/2.0>), which permits unrestricted use, distribution, and reproduction in any medium, provided the original work is properly cited.

The license is subject to the Beilstein Journal of Organic Chemistry terms and conditions: (<http://www.beilstein-journals.org/bjoc>)

The definitive version of this article is the electronic one which can be found at: [doi:10.3762/bjoc.4.22](https://doi.org/10.3762/bjoc.4.22)